



Assessment grid			
Subject: <b>Science</b>	Year: <b>7</b>	Topic/module: <b>Acids and Alkalis</b>	
<b>KS4 target direction</b>	<b>4</b>	<b>6</b>	<b>8(9)</b>
<b>Advanced</b>	Enrichment/extension – reaching, or part of, next pathway → Features of work may include:	Enrichment/extension – reaching, or part of, next pathway → Features of work may include:	Enrichment/extension Features of work may include:
<b>Secure</b> <i>Students must achieve competence in <b>all</b> statements before being judged 'Secure'</i>	<b>Secure</b> The student can: <ul style="list-style-type: none"><li>• Name some common properties of acids and alkalis</li><li>• State that concentrated acids are more harmful than dilute acid</li><li>• Describe, in simple terms, what the key words 'concentrated' and 'dilute' mean.</li><li>• State that indicators will be different colours in acids, alkalis, and neutral solutions</li><li>• State simply what happens during a neutralisation reaction</li><li>• Give one example of a neutralisation reaction</li><li>• State the type of chemical made when an acid and alkali react.</li><li>• Match the type of salt that will form from the type of acid used.</li></ul>	<b>Secure</b> The student can: <ul style="list-style-type: none"><li>• Compare the properties of acids and alkalis</li><li>• Describe the differences between concentrated and dilute solutions of an acid</li><li>• Use the pH scale to measure acidity and alkalinity</li><li>• Describe how pH changes during neutralisation reactions</li><li>• State examples of useful neutralisation reactions</li><li>• Describe what a salt is</li><li>• Predict the salts formed when acids react with metals or bases</li></ul>	<b>Secure</b> The student can: <ul style="list-style-type: none"><li>• Compare the different particles found in acids and alkalis.</li><li>• Explain what 'concentrated' and 'dilute' mean, in terms of the numbers of particles present</li><li>• Use a variety of indicators to measure acidity and alkalinity and explain how they work.</li><li>• Interpret a graph of pH changes during a neutralisation reaction.</li><li>• Explain why neutralisation reactions are useful in the context of specific examples</li><li>• Predict the formulae for products of reactions between acids and metals, or acids and bases.</li></ul>
<b>Developing</b>	Mostly secure – one or more gaps For example:	Mostly secure – one or more gaps For example:	Mostly secure – one or more gaps For example:
<b>Beginning</b>	Significant gaps	Significant gaps	Significant gaps